

# Determination of formaldehyde by polarography

### Summary

Formaldehyde can be determined reductively at the DME. Depending on the sample composition it can be possible to determine the formaldehyde directly in the sample. If interferences occur, a sample preparation may be necessary, i.e. absorption, extraction, or distillation.

Two methods are described. In the first method formaldehyde is reduced directly in alkaline solution. Higher concentrations of alkaline or alkaline earth metals interfere. In those cases the second method can be applied. Formaldehyde will be derivatized with hydrazine forming the hydrazone, which can be measured polarographically in acidic solution.

### Instruments

VA instrument capable of operating a Multi-Mode Electrode and supporting differential pulse (DP) measuring mode

### Electrodes

WE	Multi-Mode Electrode pro	6.1246.120
	Mercury drop capillary	6.1226.030
RE	Ag/AgCl reference electrode	6.0728.x20
	Ag/AgCl/KCl (3 mol/L)	
	Electrolyte vessel Filled with c(KCl) = 3 mol/L	6.1245.010
AE	Pt rod electrode	6.0343.x00

### Sample preparation

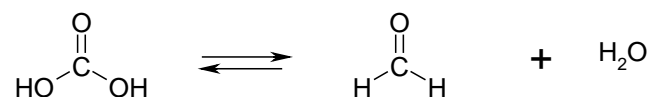
- Waste water, solutions, and plating baths (e.g. electroless copper baths) can be analyzed directly.
- Plastics and textiles are ground and extracted for approximately 20 min in a shaker with c(LiOH) = 0.05 mol/L. The extract is then separated by filtration or centrifugation.
- Air samples are absorbed in c(LiOH) = 0.05 mol/L: 2 to 10 L air with a flow rate of 500 mL/min.
- Fish, crustacean, and film material are ground. 1 to 5 g of the sample are suspended in 20 mL distilled water, 1 mL w(H<sub>2</sub>SO<sub>4</sub>) = 30 % is added. The formaldehyde is then steam distilled and absorbed in c(LiOH) = 0.05 mol/L.

## Method 1: Determination of formaldehyde in alkaline solution

### Theory

Formaldehyde can be reduced directly to methanol in alkaline solution. This reaction is used to determine formaldehyde polarographically. The method is suitable for samples, which do not contain a too high content of alkaline or earth alkaline metal ions.

A chemical equilibrium precedes the electrochemical reaction:



### Reagents

All of the used reagents must be of analysis quality (for analysis).

- Lithium hydroxide monohydrate (LiOH · H<sub>2</sub>O), for analysis, CAS 1310-66-3
- Ethylenediaminetetraacetic acid (EDTA, Titriplex II®, Komplexon II®, Idranal II®), for analysis, CAS 60-00-4
- Formaldehyde solution, w(HCHO) = 37%, for analysis, CAS 50-00-0
- Ultrapure water, resistivity >18 MΩ·cm (25 °C), type I grade (ASTM D1193)

### Solutions

Absorption solution	c(LiOH) = 0.05 mol/L 2.1 g LiOH are dissolved in 1000 mL water.
Supporting electrolyte	c(LiOH) = 0.22 mol/L, c(EDTA) = 0.02 mol/L 9.23 g LiOH and 5.85 g EDTA are dissolved in 1000 mL water.

### Standard solutions

Formaldehyde standard solution	$\beta(\text{HCHO}) = 200 \text{ mg/L}$ The standard solution is prepared from a concentrated formaldehyde solution, which is diluted with water. The exact concentration is determined by titration.
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### Analysis

5 mL sample or absorption solution are pipetted into the measuring vessel and 5 mL supporting electrolyte are added. The measuring solution is purged for 600 s.

### Measuring solution

5 mL supporting electrolyte

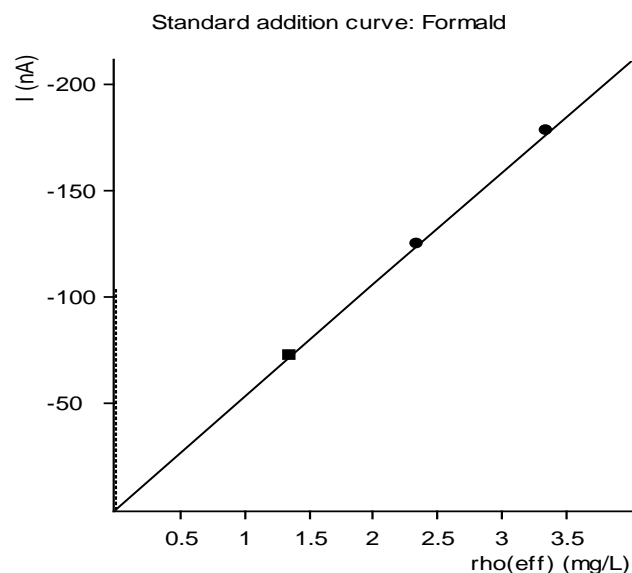
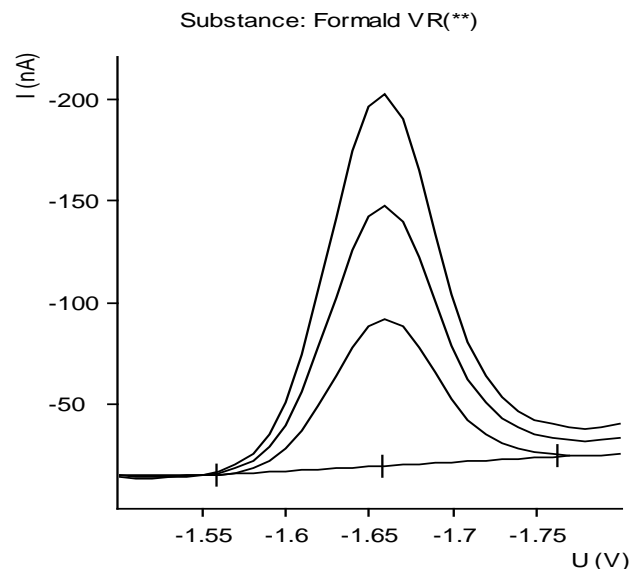
5 mL sample or absorption solution

The concentration of formaldehyde is determined by standard addition technique.

### Parameters

Voltammetric	
Electrode operating mode	DME
Measuring mode	DP – Differential pulse
Stirring rate	$2000 \text{ min}^{-1}$
Equilibration time	5 s
Sweep	
Start potential	-1.4 V
End potential	-1.8 V
Potential step	0.01 V
Potential step time	0.4 s
Sweep rate	0.025 V/s
Pulse amplitude	0.05 V
Substance	
Name	Formaldehyde
Characteristic potential	-1.65 V

### Example



### Result

Sample size	5.0 mL
$\beta(\text{Formaldehyde})$	2.7 mg/L

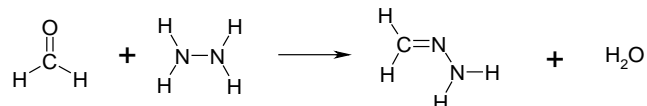
### Comments

- The linear range is between  $300 \mu\text{g/L}$  and  $300 \text{ mg/L}$  formaldehyde in the sample. Higher concentrations have to be diluted.
- The formaldehyde peak is very close to the sodium peak. It is important to avoid high concentrations of sodium ions in the measurement solution.
- If the formaldehyde peak is not separated from the rising background current (sodium peak), we recommend to use method 2.

## Method 2: Determination of formaldehyde as hydrazone

### Theory

In acidic solutions formaldehyde and hydrazine condense to a hydrazone. This can be determined polarographically.



### Reagents

All of the used reagents must be of analysis quality (for analysis).

- Sulfuric acid,  $w(\text{H}_2\text{SO}_4) = 96\%$ , for analysis, CAS 7664-93-9
- Hydrazine sulfate, for analysis, CAS 10034-93-2
- Formaldehyde solution,  $w(\text{HCHO}) = 37\%$ , for analysis, CAS 50-00-0
- Ultrapure water, resistivity  $>18 \text{ M}\Omega \cdot \text{cm}$  (25 °C), type I grade (ASTM D1193)

### Solutions

The standard solutions are prepared from a concentrated formaldehyde solution, which is diluted with water. The exact concentration is determined by titration.

Hydrazine solution  $\beta(\text{hydrazine sulfate}) = 20 \text{ g/L}$   
2 g hydrazine sulfate are dissolved in 100 mL water

### Standard solutions

Formaldehyde standard solution  $\beta(\text{HCHO}) = 100 \text{ mg/L}$   
The standard solution is prepared from a concentrated formaldehyde solution, which is diluted with water. The exact concentration is determined by titration.

### Analysis

10 mL (diluted) sample or adsorption solution are pipetted into the measuring vessel. 50  $\mu\text{L}$  sulfuric acid and 1 mL hydrazine solution are added. The measuring solution is purged for 300 s.

### Measuring solution

10 mL (diluted) sample or absorption solution

50  $\mu\text{L}$  sulfuric acid

1 mL hydrazine solution

The concentration of formaldehyde is determined by standard addition technique.

### Parameters

#### Voltammetric

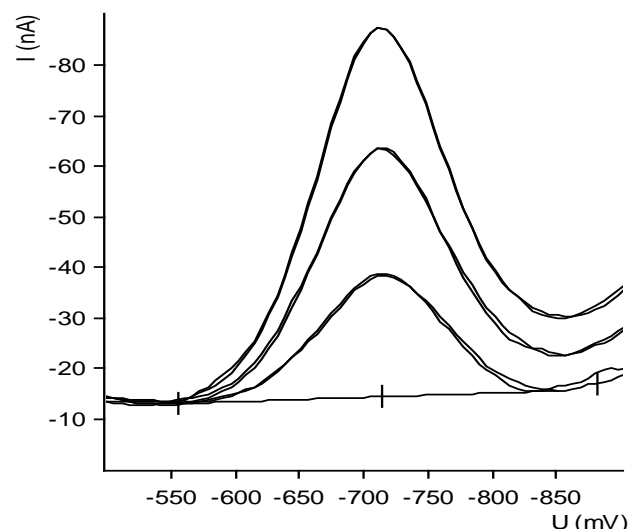
Electrode operating mode	DME
Measuring mode	DP – Differential pulse
Stirring rate	2000 $\text{min}^{-1}$
Equilibration time	10 s
Sweep	
Start potential	-0.5 V
End potential	-1.2 V
Potential step	0.006 V
Potential step time	0.3 s
Sweep rate	0.02 V/s
Pulse amplitude	0.05 V

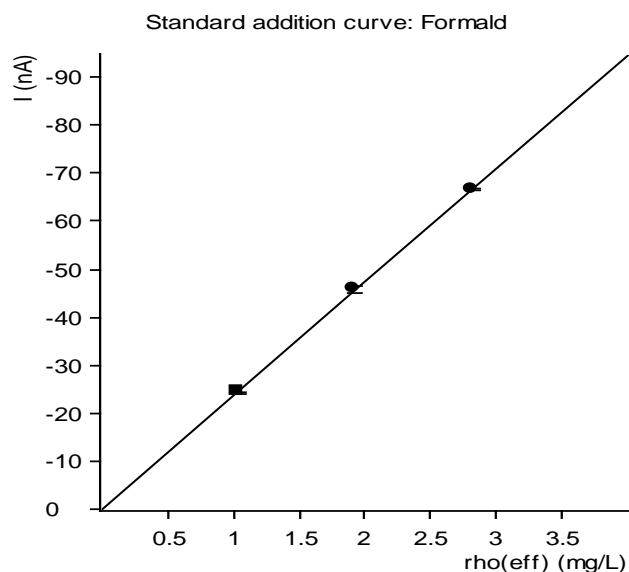
#### Substance

Name	Formaldehyde
Characteristic potential	-0.72 V

### Example

Substance: Formald VR(\*\*)





### Result

Sample size	0.1 mL
$\beta$ (Formaldehyde)	114.9 mg/L

### Comments

- The linear range is between approximately 50  $\mu\text{g/L}$  and 30 mg/L. Higher concentrations have to be diluted.
- This method is recommended, if interferences by higher contents of sodium ions occur when using method 1 in alkaline solutions.

### References

- Absenger R., Schliefer K., Analysis of textile-finishing agents. XI. Polarographic determination of free formaldehyde in nitrogen-containing cross-linking agents, *Textilveredelung* 8 (1973) 580-581
- Ardelt H.W., Opel P.H., Steindorf W., Polarographische Bestimmung des freien Formaldehyds in Kunststoffbedarfsgegenständen, *Pharmazie* 16(1961) 128-131, Ref.: *Fresenius Z. Anal. Chem.* 186 (1962) 457
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- Linhart K., Bestimmung von freiem und gebundenem Formaldehyd in Textilhilfsmitteln durch Wechselstrompolarographie, *Melliand Textilber.* 56 (1975) 240-245
- Novak V., Die polarographische Bestimmung von Verunreinigungen in Monomeren. II. Formaldehyd- und Benzaldehydbestimmung in Styrol, *Chem Prumysl.* 19(1969) 25-26, Ref.: *Fresenius Z. Anal. Chem.* 250 (1970)
- Sandler S., Chung Y.H., Polarographic determination of hydrogen peroxide, formaldehyde and acetaldehyde in mixtures, *Anal. Chem.* 30 (1958) 1252-1255
- Wild P.W., Einige polarographische Methoden in der Galvanotechnik. Formaldehyd in einem Reduktiv-Kupfer-Bad, *Galvanotechnik* 62 (1971) 2
- Zaitseva Z.V., Prokhorova E.K., Polarographic determination of formaldehyde in production area air, *Zh Anal Khim* 40 (1985) 1308-1311, Ref.: *Fresenius Z. Anal. Chem.* 324 (1986)
- Metrohm Info 2/98, 22-23

# Appendix

## Report for the example determination of formaldehyde according to method 1

```

===== METROHM 746 VA TRACE ANALYZER (5.746.0101) =====
Determ.      : 05071509      User:      Date: 1999-05-07
Modified    : 2000-12-01 16:35:19  Run : 0      Time: 15:09:05
Sample table: -
  
```

Pos.	Ident.1/S1 sample	Ident.2/S2	Ident.3/S3	Method.call	Sample size/S0 5 mL
-----					
Method	: AB196_1a				
Title	: Determination of Formaldehyde. AB196				
Remark1	: 5 mL electrolyte + 5 mL sample				
Remark2	:				

Substance	Formald	Mass conc.:	2.725 mg/L	Mass	: 13.62 ug	Comments
MC.dev.	:	0.050 mg/L (1.83%)		Add.mass	: 10 ug	
Cal.dev.	:	-		V0.sample:	5 mL	

VR	U/mV	I/nA	I.mean	Std.dev.	I.delta	Comments
00	-1658	-71.87	-71.87			
10	-1657	-123.4	-123.4		-51.52	
20	-1656	-175.4	-175.4		-52.00	

Substance	Techn.	Y.reg/offset	Slope	Nonlin.	Mean deviat.
Formald	std.add.	-7.171e-08	-5.264e-05		4.059e-10

Final results	Formald =	2.7246 mg/L	+/-	Res.dev.	0.050	%	1.83	Comments
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## Method print for the determination of formaldehyde according to method 1

```

===== METROHM 746 VA TRACE ANALYZER (5.746.0101) =====
Method: AB196_1a.mth      OPERATION SEQUENCE
Title : Determination of Formaldehyde. AB196
  
```

	Instructions	t/s	Main parameters	Auxiliary parameters
1	DOS/M		V.added 5.000 mL	
2	REM		electrolyte	
3	SMPL/M		V.fraction mL	V.total L
4	STIR		Rot.speed 2000 /min	
5	PURGE	600.0		
6	(ADD			
7	NOP	10.0		
8	SEGMENT		Segm.name pol	
9	ADD>M		Soln.name For-Std	V.add 0.050 mL
10	ADD)2			
11	END			

```

Method: AB196_1a      SEGMENT
                        pol
  
```

	Instructions	t/s	Main parameters	Auxiliary parameters
1	0PURGE			
2	0STIR	5.0		
3	(REP			
4	DME			
5	DPMODE		U.ampl -50 mV	t.meas 30.0 ms
			t.step 0.40 s	t.pulse 40.0 ms
6	SWEEP	17.6	U.start -1400 mV	U.step 10 mV
			U.end -1800 mV	Sweep rate 25 mV/s
7	REP)0			
8	PURGE			
9	STIR		Rot.speed 2000 /min	
10	OMEAS		U.standby mV	
11	END			

```

Method: AB196_1a      SUBSTANCES
                        Formald - pol
  
```

Recognition		Display / Plot	
U.verify	-1650 mV	I.scale	auto
U.tol (+/-)	50 mV	U.div	50.00 mV/cm
U.width min	10 mV	U.begin	-1500 mV
U.width max	400 mV	U.end	-1800 mV
I.threshold	200 pA		

Baseline		Evaluation	
Type	linear	Mode	VA
Scope	whole	Quantity	I.peak
dU.front	auto	Sign. digits	4
S.front	auto		
dU.rear	auto		
S.rear	auto		

Calibration	2000-12-01 16:38:17	Coefficients	
Technique	std.add.	Y.reg	-7.171e-08
Curve type	linear	Slope	-5.264e-05
		Nonlin.	
		Mean dev.	4.059e-10

Additions

Soln.name	For-Std			
Mass conc.	200 mg/L	g/L	g/L	g/L
Range min	0 g/L	g/L	g/L	g/L
Range max	g/L	g/L	g/L	g/L
M.conc./cm	g/L	g/L	g/L	g/L

Method: AB196\_1a                      CALCULATION  
 max. 15 lines

Quantity	Formula (R##, C##, A##)	Res.unit	Sig.dig.
Formald	R1000=MC:Formald	#g/L	5

**Report for the example determination of formaldehyde according to method 2**

===== METROHM 746 VA TRACE ANALYZER (5.746.0101) =====  
 Determin. : 06211516                      User:                      Date: 1994-06-21  
 Modified : 2000-12-01 17:00:42        Run : 0                      Time: 15:16:55  
 Sample table: -

Pos.	Ident.1/S1	Ident.2/S2	Ident.3/S3	Method.call	Sample size/S0
	Std	For			100 uL

Method : AB196\_2  
 Title : Determination of Formaldehyde. AB196 method 2  
 Remark1 : 10 ml Wasser + 50 µl H2SO4 + 1 ml Hydrazinsulfat  
 Remark2 : + Probe

Substance : Formald	Comments
Mass conc.: 114.9 mg/L	Mass : 11.49 ug
MC.dev. : 2.58 mg/L (2.25%)	Add.mass : 10 ug
Cal.dev. : -	V0.sample: 100 uL

VR	U/mV	I/nA	I.mean	Std.dev.	I.delta	Comments
00	-715	-24.37	-24.20	0.2360		rear overlapping
01	-717	-24.04				
10	-714	-45.83	-45.39	0.6269	-21.18	rear overlapping
11	-712	-44.95				
20	-711	-65.50	-65.39	0.1461	-20.00	rear overlapping
21	-711	-65.29				

Substance	Techn.	Y.reg/offset	Slope	Nonlin.	Mean deviat.
Formald	std.add.	-2.434e-08	-2.362e-05		4.193e-10

Final results	+/-	Res.dev.	%	Comments
Formald = 114.90 mg/L		2.58	2.25	

**Method print for the determination of formaldehyde according to method 2**

```

===== METROHM 746 VA TRACE ANALYZER (5.746.0101) =====
Method: AB196_2 .mth          OPERATION SEQUENCE
Title : Determination of Formaldehyde. AB196 method 2
    
```

	Instructions	t/s	Main parameters	Auxiliary parameters	
1	DOS/M		V.added 11.050 mL		
2	REM		Wasser/Hydr./Puffer		
3	STIR		Rot.speed 2000 /min		
4	PURGE	300.0			
5	SMPL>M		V.fraction mL	V.total	L
6	(ADD				
7	NOP	10.0			
8	SEGMENT		Segm.name pol		
9	ADD>M		Soln.name For-Std	V.add	0.100 mL
10	ADD)2				
11	END				

```

Method: AB196_2          SEGMENT
                        pol
    
```

	Instructions	t/s	Main parameters	Auxiliary parameters	
1	0PURGE				
2	0STIR	10.0			
3	(REP				
4	DME				
5	DPMODE		U.ampl -50 mV	t.meas	20.0 ms
			t.step 0.30 s	t.pulse	40.0 ms
6	SWEEP	36.0	U.start -500 mV	U.step	6 mV
			U.end -1200 mV	Sweep rate	20 mV/s
7	0MEAS		U.standby mV		
8	REP)1				
9	PURGE				
10	STIR		Rot.speed 2000 /min		
11	END				

```

Method: AB196_2          SUBSTANCES
                        Formald - pol
    
```

Recognition		Display / Plot	
U.verify	-720 mV	I.scale	auto
U.tol (+/-)	30 mV	U.div	50.00 mV/cm
U.width min	10 mV	U.begin	-500 mV
U.width max	200 mV	U.end	-900 mV
I.threshold	250 pA		

Baseline		Evaluation	
Type	linear	Mode	VA
Scope	whole	Quantity	I.peak
dU.front	auto	Sign. digits	4
S.front	auto		
dU.rear	auto		
S.rear	auto		

Calibration		Coefficients	
Calibration	2000-12-01 17:01:39	Y.reg	-2.434e-08
Technique	std.add.	Slope	-2.362e-05
Curve type	linear	Nonlin.	
		Mean dev.	4.193e-10

**Additions**

Soln.name	For-Std			
Mass conc.	100 mg/L	g/L	g/L	g/L
Range min	g/L	g/L	g/L	g/L
Range max	g/L	g/L	g/L	g/L
M.conc./cm	g/L	g/L	g/L	g/L

```

Method: AB196_2          CALCULATION
                        max. 15 lines
    
```

Quantity	Formula (R##, C##, A##)	Res.unit	Sig.dig.
Formald	R1000=MC:Formald	#g/L	5